

Valency of Common Monatomic and Polyatomic Ions

		Positive Ions				Negative Ions		
		+1	+2	+3	+4	-1	-2	-3
Monatomic Ions	H ⁺ hydrogen	Be ²⁺ beryllium	Al ³⁺ aluminium	Pb ⁴⁺ lead (IV)	F ⁻ fluoride	O ²⁻ oxide	N ³⁻ nitride	
	Li ⁺ lithium	Mg ²⁺ magnesium	Co ³⁺ cobalt (III)	Sn ⁴⁺ tin (IV)	Cl ⁻ chloride	S ²⁻ sulfide	P ³⁻ phosphide	
	Na ⁺ sodium	Ca ²⁺ calcium	Cr ³⁺ chromium (III)		Br ⁻ bromide			
	K ⁺ potassium	Sr ²⁺ strontium	Fe ³⁺ iron (III)		I ⁻ iodide			
	Ag ⁺ silver	Ba ²⁺ barium	Au ³⁺ gold (III)					
	Cu ⁺ copper (I)	Ni ²⁺ nickel	Mn ³⁺ manganese (III)					
	Au ⁺ gold (I)	Cd ²⁺ cadmium						
		Zn ²⁺ zinc						
		Cr ²⁺ chromium (II)						
		Co ²⁺ cobalt (II)						
		Cu ²⁺ copper (II)						
		Fe ²⁺ iron (II)						
		Mn ²⁺ manganese (II)						
		Pb ²⁺ lead (II)						
	Sn ²⁺ tin (II)							
Polyatomic Ions	NH ₄ ⁺ ammonium				CH ₃ COO ⁻ acetate	CO ₃ ²⁻ carbonate	PO ₄ ³⁻ phosphate	
					HCO ₃ ⁻ hydrogen carbonate	SO ₄ ²⁻ sulfate		
					OH ⁻ hydroxide			
					NO ₃ ⁻ nitrate			